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## Atomic structure practice 2 worksheet answer key

Which of the following terms refers to the smallest part of a compound that has all the properties of that compound? Every atom (except the hydronium ion) has two main components, a NUCLEUS with PROTONS and NEUTRONS and one or more ORBITS (or SHELLS) that contain one or more electrons. An atom is composed of protons, neutrons, and electrons. The protons and neutrons are found in the nucleus. The electrons are found in the electron cloud, which is an area that surrounds the nucleus. What is the mass number, symbol, and charge of an ion that contains 35 protons, 45 neutrons and 36 electrons? Write the complete chemical symbol for the ion with 82 protons and 80 electrons. Complete this here chart for us. Show us where the electrons are located. Write the complete electron configuration for each of the following elements. Identify the element that corresponds to each of the following electron configurations? Draw orbital diagrams for the following choices. Given an element and a mass number you will tell us the number of protons, electrons, and neutrons. You will also diagram the electron configuration in an orbital diagram. The electrons found in the outer most energy level are collectively referred to as valence electrons. Lewis Dot Diagrams are used to indicate the number of valence electrons and provide us with a quick form of short hand. Given an element, like Magnesium (Atomic Number 12, Mass Number 24), please provide all the following information for this element. For identified element identify the ground state electron configuration, orbital diagram, Lewis dot diagram, and number of valence. Place a check in the correct box to indicate the classification of each form of matter. Place a check in the correct box to indicate the classification of each form of matter. Place a check to identify the form of radiation demonstrated by each reaction below. Predict the products of the following nuclear reactions. What is going on with these nuclear reactions? Radioactive substances decay at a constant rate. There are a number of unstable nuclei that decay in a given time. The time that it takes for half of the atoms in a given sample of an element to decay is termed the half-life. For questions 1 and 2 predict the missing reactant or product and identify the form of radiation demonstrated in each problem. Write the electron configuration, orbital diagram, total number of valence electrons, and Lewis dot structure. Indicate if the item presented is an element, compound, or mixture. A 75 gram sample of a radioisotope decayed to 9.375 grams in 66 years. What is the half-life? Identify the element that corresponds to each of the following electron configurations. Lewis Dot diagram, total number of valence electrons, and Lewis dot structure. Indicate if the item presented is an element, compound, or mixture. This may look a bit different than others. Charge it all up. Using the model key for each element, write the chemical formula for each molecule. Using the model key for each element, draw the model of each molecule. What is missing in here? Find those different elements and parts. The element you want to see. The number of protons is the atomic number and the mass number is the sum of the protons and neutrons. The electrons in a neutral atom equal the number of protons. Bohr models (or Bohr diagram) are diagrams that show the number of protons and neutrons in the nucleus and the electrons in each electron energy level around it. Provide the number of atoms in each molecule of the substances represented by the chemical formulas below. Which of the following terms refers to the smallest part of a compound that has all the properties of that compound? Every atom (except the hydronium ion) has two main components, a NUCLEUS with PROTONS and NEUTRONS and one or more ORBITS (or SHELLS) that contain one or more electrons. An atom is composed of protons, neutrons, and electrons. The protons and neutrons are found in the nucleus. The electrons are found in the electron cloud, which is an area that surrounds the nucleus. What is the mass number, symbol, and charge of an ion that contains 35 protons, 45 neutrons and 36 electrons? Write the complete chemical symbol for the ion with 82 protons and 80 electrons. Complete this here chart for us. Show us where the electrons are located. Write the complete electron configuration for each of the following elements. Identify the element that corresponds to each of the following electron configurations? Draw orbital diagrams for the following choices. Given an element and a mass number you will tell us the number of protons, electrons, and neutrons. You will also diagram the electron configuration in an orbital diagram. The electrons found in the outer most energy level are collectively referred to as valence electrons. 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The number of protons is the atomic number and the mass number is the sum of the protons and neutrons. The electrons in a neutral atom equal the number of protons. Bohr models (or Bohr diagram) are diagrams that show the number of protons and neutrons in the nucleus and the electrons in each electron energy level around it. Provide the number of atoms in each molecule of the substances represented by the chemical formulas below.





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