Conjugate acid of hso3

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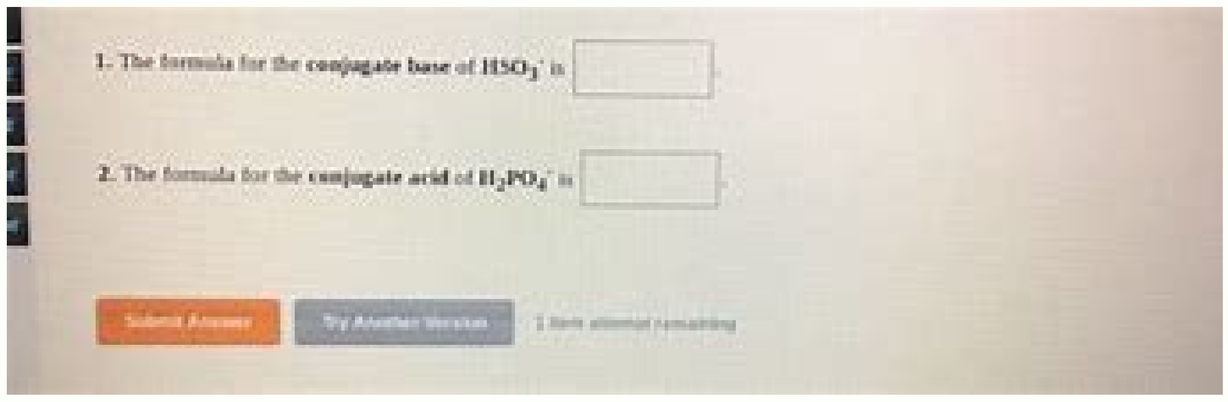
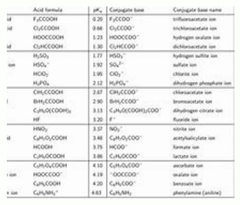




	Acid formula	pK_a	Conjugate base	Conjugate base name
ul	F_3CCOOH	0.20	F_3CCOO^-	trifluoroacetate ion
ul	Cl_3CCOOH	0.66	Cl_3CCOO^-	trichloroacetate ion
	$HOOCCOOH$	1.23	$HOOCCOO^-$	hydrogen oxalate ion
ul	$Cl_2HCCOOH$	1.30	Cl_2HCCOO^-	dichloroacetate ion
	H_2SO_3	1.77	HSO_3^-	hydrogen sulfite ion
ion	HSO_4^-	1.92	SO_4^{2-}	sulfate ion
	$HClO_2$	1.95	ClO_2^-	chlorite ion
	H_3PO_4	2.12	$H_2PO_4^-$	dihydrogen phosphate ion
	ClH_2CCOOH	2.87	ClH_2CCOO^-	chloroacetate ion
	BrH_2CCOOH	2.90	BrH_2CCOO^-	bromoacetate ion
	$C_3H_4O(COOH)_3$	3.13	$C_3H_4O(COOH)_2COO^-$	dihydrogen citrate ion
	HF	3.20	F^-	fluoride ion
	HNO_2	3.37	NO_2^-	nitrite ion
d	$C_6H_5O_2COOH$	3.46	$C_6H_5O_2COO^-$	acetylsalicylate ion
	$HCOOH$	3.75	$HCOO^-$	formate ion
	C_3H_6OCOOH	3.86	$C_3H_6OCOO^-$	lactate ion
	$C_6H_7O_6COOH$	4.10	$C_6H_7O_6COO^-$	ascorbate ion
ion	$HOOCCOO^-$	4.19	$^-OOCCOO^-$	oxalate ion
	C_6H_5COOH	4.20	$C_6H_5COO^-$	benzoate ion
+ ion	$C_6H_5NH_3^+$	4.63	$C_6H_5NH_2$	phenylamine (aniline)

Sulfur dioxide is used as a preservative in wine maki

The following equations describe how sulfur dioxide

$H_2O + SO_2 \rightleftharpoons HSC$

$HSO_3^- + H^+ \rightleftharpoons SO_3$

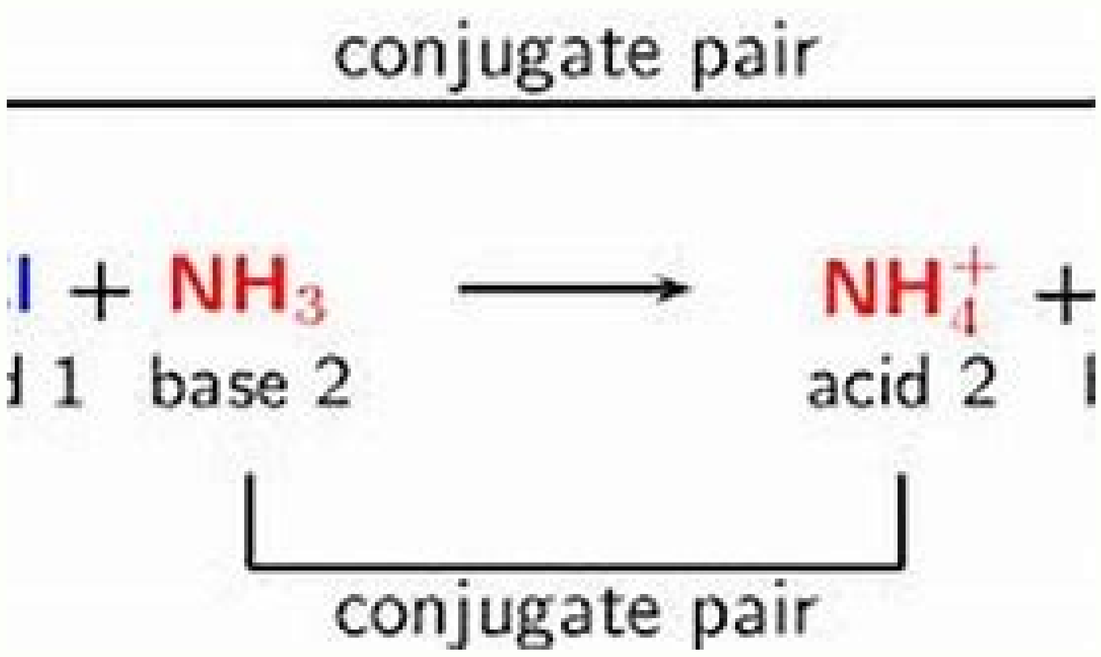
Which statement about these two reactions is correc

A HSO_3^- acts as a base.

B SO_2 acts as an oxidising agent.

C SO_3^{2-} acts as an acid.

D SO_3^{2-} acts as a reducing agent.



1. The formula for the **conjugate base** of HSO_3^- is .

2. The formula for the **conjugate acid** of HCO_3^- is .

Conjugate acid and base of hso3-. Formula of the conjugate acid of hso3−. Identify the conjugate acid of hso3-. What is the conjugate acid of hso3−hso3−. Conjugate acid of hso3-. What is the conjugate acid of hso3- quizlet. Conjugate acid of hso3-1. Give the conjugate acid of hso3-.

Not to be confused with hydrosulfite. Equilibrium reaction that connects the two bisulfite tautumers. The bisulphite ion (IUPAC recommended nomenclature: hydrogen sulphite) is the HSOÃ¢3 ion. Salts containing the HSOÃ¢ ÌÀ3 ion are also known as Â"bleach sulphitesÂ".[1] Sodium bisulfite is used interchangeably with sodium metabisulfite (Na2S2O5). Sodium metabisulfite is dissolved in water to give a solution of Na+HSOÃ¢ ÌÀ3. Na2S2O5 + H2O 2 Na[HSO3] Structure Bisulfite Anion exists in solution as a mixture of two tautomers. A tautomer has the proton attached to one of the three oxygen centers. In the second tautomer the proton resides in the sulfur. The S-protonated tautomer has C3v symmetry. The O-protonated tautomer has only Cs symmetry. Reactions Tautomerization There are two tautomers of bisulfite. They interconvert easily, but can be characterized individually by various spectroscopic methods. The following have been observed by 17O NMR spectroscopy:[1][2] HSO3Ã¢ Ã¢Â½ Ã¢ Ã¢ Ã¢ {\displaystyle {\ce {>}} SO2(OH)Ã¢ K = 4.2 Acid-base reactions Bisulfite solutions are typically prepared by treatment of sulfur dioxide with an aqueous base:[3] SO2 + OHÃ¢ Ã¢ HSOÃ¢3 HSOÃ¢3 is the conjugate base of sulfurous acid, (H2SO3). HSOÃ¢3 is a weak acidic species with a pKa of 6.97. Its conjugate base is sulfite, SO2Ã¢3: HSOÃ¢3 Ã¢ SO2Ã¢3 + H+ Dehydration/hydration equilibrium The attempt to isolate the common salts of bisulfite results in the dehydration of the anion with the formation of metabisulfite (S2O2Ã¢5), also known as disulfite: 2 HSOÃ¢3 Ã¢ S2O2Ã¢5 + H2O Due to this balance, it is not possible to obtain the anhydrous salts of sodium and potassium from the bisulfite. However, there are some reports of anhydrous bisulfites with large contra-ions[4] Structure of disulfite ion (also known as metabisulfite).[5] Reactions Redox Bisulfite is a good reducing agent, especially for oxygen depuration: 2 HSOÃ ¢ 3 + O2 Ã ¢ Ã ¢ Â |2 SO2Ã ¢ Â |4 + 2 H + Its reducing properties are exploited to precipitate gold from the Áuric acid (gold (gold in Aqua Regia) and reduce chromium (VI) to chromium (III). In water chlorination, sodium bisulfite is used to reduce the residual 'chlorine' that can have a negative impact on aquatic life. Organic synthesis in organic chemistry, "sodium bisulfite" is used to form adducts with aldehyde and with certain cyclic ketones. These adducts are Î±-hydroxy-sulfonic. [6] This reaction is useful for the purification of aldehydes. Bisulfite adducts precipitate as solution solids. The reaction can be reversed at the base. [7] Examples of such procedures are described for Benzaldehydo, [8] 2-tralone, â[9] citral, [10] the ethyl ester of pyrovic acid [11] and glloxal. [12] In the cyclohexanone ring expansion reaction with diazald, it is reported that the reaction of the bisulfite allows the differentiation between the primary reaction product cycloheptanone and the main contaminating cyclooctanone. [13] Another use of bisulfite in organic chemistry is as a mild reducing agent, e.g. to remove traces or excess amounts of chlorine, bromine, iodine, hypochlorite salts, osmatate esters, chromium trioxide and potassium permanganate. Sodium bisulfite is a discoloration agent in purification procedures because it reduces strongly coloured oxidizing agents, conjugated alkenes and carbonyl compounds. Bisulfite is also the key ingredient in the Bucherer reaction. In this reaction, an aromatic hydroxyl group becomes the corresponding amine group. This is a reversible reaction. The first step in this reaction is an addition reaction of sodium bisulfite to an aromatic double bond. Bucherer Carbazole synthesis is a related organic reaction that uses sodium bisulfite as a reagent. Bisulfite DNA sequencing More information: Bisulfite that sequenced the chemical reaction underlying the cytosine bisulfite-mediated conversion to Sodium bisulfite is used in the analysis of the methylation state of cytosines in DNA. In this technique, sodium bisulfite deamines cytosine into uracil, uracil, does not affect 5-methylcytosine, a form methylated cytosine with a methyl group attached to the carbon 5. When the treated DNA with bisulfite was amplified AAV chain reaction polymerase, uracil is amplified as thymine and methylated cytosine are amplified as cytosine. ATs © techniques of DNA sequencing is then used to read the sequence of the bisulfite treated DNA. Cytosines that are read as cytosine © s after the sequencing represent methylated cytosines, while those read as thymine represent unmethylated cytosines in genomic DNA. [14] References ^ b Greenwood, Norman N .; Earnshaw, Alan (1997). Chemistry of the Elements (2nd ed.). Butterworth man. ISBN 978-0-08-037 941-8. Horner, D. A .; Connick, R. E. (1986). Â "balance ratio for the isomerization of bisulfite ion HSOÃ ¢ Î Î ¢ 3 SO3HÃ .Â» QuÃmica inorganic. 25 (14): 2414 Â "2417. doi: 10.1021 / ic00 234a026. â BarberÃ¡, Joseph © JimÃ © nez; Metzger, Adolf; Wolf, Manfred (2000). Â «sulfites, thiosulfates, and Dithionitesl QuÃmica.Â" The Encyclopedia of Industrial Chemistry Ullmann. Weinheim: Wiley-Vch. doi: 10.1002 / 14356 007.a25_477. â Maylor, R .; Gill, J. B .; Goodall, D.C. (July 1971). Â "Some studies on cobalt sulfite anhidro.Â" Journal of Inorganic and Nuclear Chemistry. 33 (7): 1975a ¢ â ²1979. doi: 10.1016 / 0022-1902 (71) 80 558-4. 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Consultado de Â« 2 Nombres de 2-tetralona Nombre preferente IUPAC 3,4-Dihydronaphthalen-2 (1H) -one Otros nombres Ã2-tetralona; Identificadores de 2-tetralona NÃºmero CAS 530-93-8Ã Y 3D model (JSmol) Imagen interactiva ChemSpider 61 564Ã N ECHA InfoCard 100.007.727 EC Number 208-498-3 PubChem CID 68 266 CompTox Dashboard (EPA) DTXSID0 060 193 InChI InChI=1S/C10H10O/c11-10-6-5-8-3-1-2-4-9 (8) 7-10/h1-4H,5-7H2Ã NKey:Ã KCKZIWSINLBROE-UHFF FAOYSA-NÃ NInChI=1/C10H10O/c11-10-6-5-8-3-1-2-4-9 (8) 7-10/h1-4H,5-7H2Key:Ã CACIWSINLBROE-UHFFFAOYAE SMILES c1ccc2c (c1) CCC (=O) Propiedades FÃ³rmula quÃmica C10H10O Masa molar 146.189Â gÂÂ·molÃ¢1 Apariencia LÃquido incoloro Densidad 1.106 g/ml Punto de fusiÃ³n 18Ã ÃÂ°C (64Ã ÃÂ°F; 291Ã K) Punto de ebulliciÃ³n 70Ã Ãâ²71Ã ÃÂ°C (158Ã ÂÂ»160Ã ÃÂ°F; 343ÃÂ»344 K) /0,25 mm Solubilidad en agua en agua bÃ¡sica Salvo indicaciÃ³n en The data is given for materials in their normal state (at 25 ° C) ° C) 100° KPA). References for infobox Chemical compound 2-ttralone is an organic chemical compound with the molecular formula C10H10O. This colorless oil is an intermediate in organic synthesis. It is a derivative of tetraline ketone, a hydrogenated derivative of naphthalene. A related compound is 1-ttralone. 2-ttralone is prepared by reductive cleavage of 2-naphthyl ethers. [1] 2-ttralone applications are an intermediate in the synthesis of a variety of pharmaceutical drugs, including L-687,384, nepinalone, napamezole, [2] spirodone and trioxifene. References ^ M. D. Soffer, M. P. Bellis, Hilda E. Gellerson, and Roberta A. Stewart "Î2-ttralone" Org. Synth. 1952, 32, 97 DOI: 10.15 227 / orgsyn.032.0097 ^ Furd Furdland, Mark P.; Bailey, Denis M.; Alexander, E. Juan; Castaldi, Michael J.; Ferrari, Richard A.; Haubrich, Dean R.; Luttinger, Daniel A.; Perrone, Mark H. (1987). "Synthesis and antidepressant properties of 4,5-dihydro-1H-dihydro-1H derivatives substituted with novels." 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